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 positive side to equal the less positive side of the half reaction $3\text{H}_2\text{O} + 2\text{Fe} + 6\text{H}^+ + 6\text{e}^- \rightarrow 2\text{O}_3 + 4\text{H} + 2 + 4\text{e}^- + 2\text{H}_2\text{O}$ put coefficients in front of molecules
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 the reactants and products this is important because a chemical equation must obey the law of conservation of mass and the law of constant proportions i.e.
 number of atoms of each element must exist on the reactant side and the product side of the equation
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 you never change subscripts you add coefficients coefficients are whole number multipliers if for example you write $2\text{H}_2\text{O}$ that means you have 2 times the n
 atoms in each water molecule which would be 4 hydrogen atoms and 2
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 was first coined or discovered by a german chemist named jeremias richter even though
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